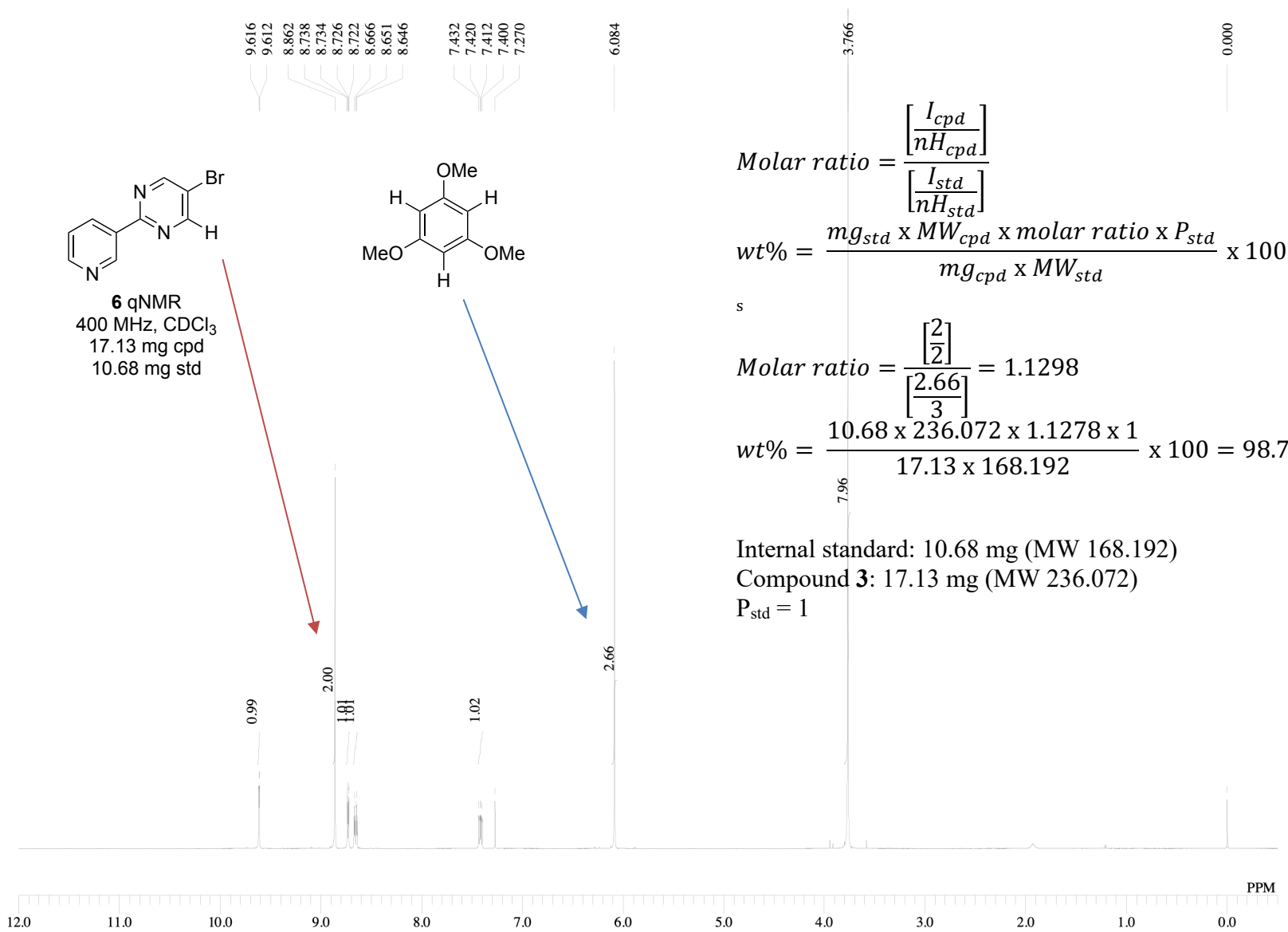




$$\text{Molar ratio} = \frac{\left[\frac{I_{cpd}}{nH_{cpd}} \right]}{\left[\frac{I_{std}}{nH_{std}} \right]}$$

$$wt\% = \frac{mg_{std} \times MW_{cpd} \times \text{molar ratio} \times P_{std}}{mg_{cpd} \times MW_{std}} \times 100$$

$$\text{Molar ratio} = \frac{\left[\frac{2}{2} \right]}{\left[\frac{2.66}{3} \right]} = 1.1298$$

$$wt\% = \frac{10.68 \times 236.072 \times 1.1278 \times 1}{17.13 \times 168.192} \times 100 = 98.7$$

Internal standard: 10.68 mg (MW 168.192)

Compound **3**: 17.13 mg (MW 236.072)

$$P_{\text{std}} = 1$$